**Name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Assignment:7**

***Review for Standard PS:9a***

1. What is an ionic bond? What is a covalent bond? How are they different?
2. What kind of bond would the following compounds form?
   1. CO2
   2. BaCl2
3. What are valence electrons?
4. Answer the following questions based on the diagram to the right:
   1. In this diagram the dots represent what?
   2. Would this be a covalent or ionic formula?
   3. Write the formula for this compound:
5. Answer the following questions based on the diagram to the right:
   1. In this diagram the dots represent what?
   2. Would this be a covalent or ionic formula?
   3. Write the formula for this compound:

1. Calculate the atomic tallies of the following formulas and find the molecular mass .

5 BaCl2 4 C6H12O6  NaNO3

1. **Determine the percent composition of the following**

BaCl2 **what percent is the Ba? the Cl?**

C6H12O6 **what percent is the C? the H? the O?**

NaNO3 **what percent is the Na? the N? the O?**

1. Write the formulas and charges, then criss cross to get the ionic compound

chromium (VI) sulfate \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

calcium bromide \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

ammonium sulfate \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

copper (II) oxide \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

platinum (IV) phosphate \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

1. Write the names for the following ionic compounds
2. AlAs \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
3. KMnO4 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
4. Cr(CN)6\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
5. SnSO3 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
6. VF5 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

1. Write the names of the covalent compounds
2. N2O5 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
3. P2O5 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
4. N2O4 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
5. CS2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
6. OF2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

11. Write the formulas of the following covalent compounds

1. antimony tribromide \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
2. hexaboron silicide \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
3. chlorine dioxide \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
4. hydrogen iodide \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
5. iodine pentafluoride \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_