

Assignment: 3

Name: PILARZ-KEY

Hour: \_\_\_\_\_

## Common Polyatomic Ions

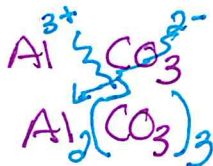
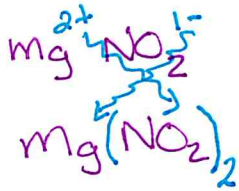
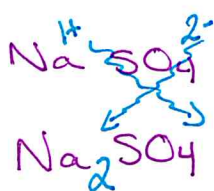
Ionic Compounds:

metal and a non-metal

- Still going to boogie or drop down

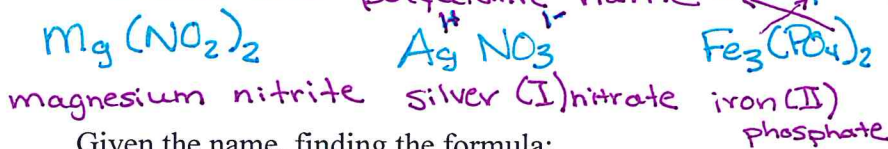
single element / polyatomic ion (more than one non-metal)

Figuring out the formula:



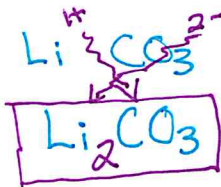
Figuring out the name:

1. Metal written first
2. Use a roman numeral in ( ) if its a transition metal
3. Use the polyatomic name

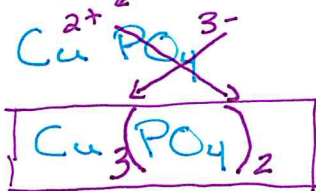


Given the name, finding the formula:

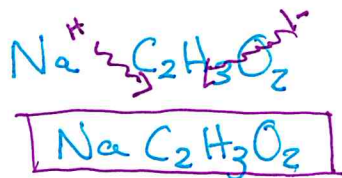
Lithium carbonate



Copper (II) phosphate



Sodium acetate



$\text{NH}_4^{1+}$	ammonium
$\text{NO}_3^{1-}$	nitrate
$\text{NO}_2^{1-}$	nitrite
$\text{ClO}^{1-}$	hypochlorite
$\text{ClO}_2^{1-}$	chlorite
$\text{ClO}_3^{1-}$	chlorate
$\text{ClO}_4^{1-}$	perchlorate
$\text{OH}^{1-}$	hydroxide
$\text{CN}^{1-}$	cyanide
$\text{C}_2\text{H}_3\text{O}_2^{1-}$	acetate
$\text{MnO}_4^{1-}$	permanganate
$\text{CO}_3^{2-}$	carbonate
$\text{HCO}_3^{1-}$	bicarbonate (hydrogen carbonate)
$\text{SO}_4^{2-}$	sulfate
$\text{SO}_3^{2-}$	sulfite
$\text{HSO}_4^{1-}$	bisulfate
$\text{HSO}_3^{1-}$	bisulfite
$\text{CrO}_4^{2-}$	chromate
$\text{Cr}_2\text{O}_7^{2-}$	dichromate
$\text{O}_2^{2-}$	peroxide
$\text{C}_2\text{O}_4^{2-}$	oxalate
$\text{PO}_4^{3-}$	phosphate
$\text{HPO}_4^{2-}$	hydrogen phosphate
$\text{H}_2\text{PO}_4^{1-}$	dihydrogen phosphate
$\text{SiO}_4^{4-}$	silicate