

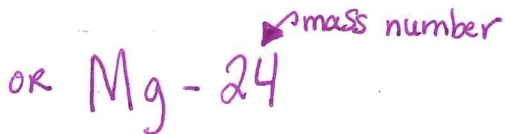
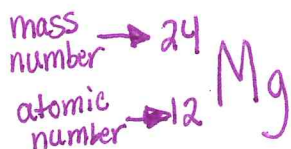
Isotopes Notes and Practice

An isotope is an atom that has the same number of protons but a different number of neutrons compared to other atoms of the same element

** If I change the number of protons, I change the element I have, so the number of protons will never change for a particular element

**The number of electrons can change, it's called an ion

We can write isotopes two ways (show them both below, be sure to label the mass number and atomic number where applicable)



*USE the mass number given instead of the one on the periodic table

Isotope Practice

Isotope	Number of Protons = atomic #	Number of Neutrons = mass number - atomic #	Number of Electrons = same as protons
Neon - 20			
Strontium - 88			
Sodium - 23			
Boron - 11			
¹⁹ F 9			
²⁷ Al 13			
⁶⁵ Zn 30			
⁴⁰ Ar 18			