

Assignment: _____

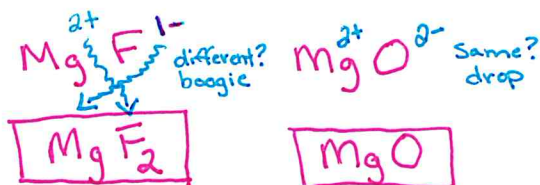
Name: Pilarz-Key

Hour: _____

Naming Ionic Compounds

Reminder:

Ionic Compounds - contain a metal and a non-metal
- metal is ALWAYS listed first



1. Alkali/Alkaline Earth Metal and a Non-metal:

1. Write out the entire name of the metal (cation)
2. Then the root of the non-metal (anion) with -ide
* on these don't worry about subscripts

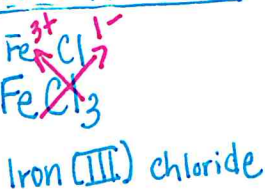
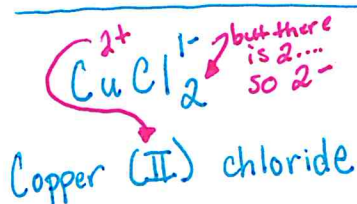
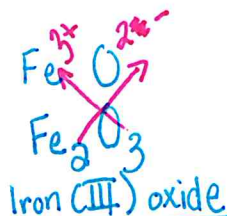
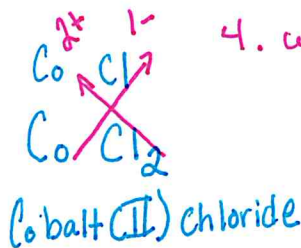
$NaCl$
Sodium chloride

CsI
cesium iodide

CaC_2
calcium carbide

2. Transition Metal and a Non-metal:

1. Write out the entire name of the transition metal
2. Figure out the charge of the transition metal (reverse criss cross method) * consider in subscripts
3. Write the charge in parentheses
4. Write the root of the non-metal with -ide



Common Nonmetallic ROOTS

H = hydr-	B = bor-
C = carb-	N = nitr-
O = ox-	F = fluor-
Si = silic-	P = phosph-
S = sulf-	Cl = chlor-
As = arsen-	Se = selen-
Br = brom-	Sb = antimony-
I = iod-	

will be part of your notebook... but turn in 1st ✓

Formulas to Names for Ionic Compounds

Name _____

The general way to name a compound is to use the first element name followed by the second element with an "ide" ending or suffix. Example: NaCl = Sodium Chloride. Some atoms of transition metals have more than one possible ion. Use a Roman numeral to indicate which one is used. +1 charge = I +2 = II +3=III etc. You will need a list of common ions to complete this handout.

Name the following ionic compounds using Roman numerals where necessary.

